

CHAPTER 6 REVIEW*Chemical Bonding***SECTION 6-5****SHORT ANSWER** Answer the following questions in the space provided.

1. Identify the major assumption of the VSEPR theory that is used to predict the shape of atoms.

2. In water, two hydrogen atoms are bonded to one oxygen atom. Why isn't water a linear molecule?

4. What two factors determine whether or not a molecule is polar?

5. Arrange the following types of attractions in order of increasing strength, with 1 being the weakest and 4 the strongest.

_____ covalent

_____ ionic

_____ dipole-dipole

6. How are dipole-dipole attractions, _____, and hydrogen bonding similar?

SECTION 6-5 continued

7. Complete the following table:

Formula	Lewis structure	Geometry	Polar
H ₂ S			
CCl ₄			
BF ₃			
H ₂ O			
PCl ₃			
BeF ₂			
SF ₆			